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Enthalpy of Formation of Magnesium Procedure

Part A:

1. Weigh a dry calorimeter, add 100ml of 1 M HCL and reweigh.
2. Obtain magnesium ribbon, weigh .24g to nearest thousandth.
3. Assemble as shown in figure.
4. Record temperature of acid to 0.1 degrease C
5. Add the magnesium and record the highest temperature of the calorimeter system.

Part B:

1. Repeat above procedure with .8g of magnesium oxide.

Part C: conduct calculations as directed in the calculations procedure on the original lab

* Calculate mass of hydraulic acid and change in temperature.
* Calculate q(soln) and q (reac)
* Calculate q(rxn ) per mole
* Write Hesses’s law expression for the chemical reactions.
* Find change in enthalpy for the reactions
* Compare with accepted values.